## New CBCS Syllabus of Shivaji University, Kolhapur As per NEP-2020

B.Sc. Part-I: Semester-I

# INORGANIC CHEMISTRY

CHEMISTRY (DSC-3A): PAPER-I

## **General Electronic Configuration of p-Block Elements**

Group	(13) III A	(14) IV A	(15) V A	(16) VI A	(17) VII A
rracteristic electronic figuration	ns² np¹	ns² np²	ns² np³	ns² np⁴	ns² np⁵
	B 2s <sup>2</sup> 2p <sup>1</sup>	C 2s <sup>2</sup> 2p <sup>2</sup>	N 2s <sup>2</sup> 2p <sup>3</sup>	O 2s <sup>2</sup> 2p <sup>4</sup>	F 2s <sup>2</sup> 2p <sup>5</sup>
	Al 3s <sup>2</sup> 3p <sup>1</sup>	Si 3s <sup>2</sup> 3p <sup>2</sup>	P 3s <sup>2</sup> 3p <sup>3</sup>	S 3s <sup>2</sup> 3p <sup>4</sup>	Cl 3s <sup>2</sup> 3p <sup>5</sup>
	Ga 4s <sup>2</sup> 4p <sup>1</sup>	Ge 4s <sup>2</sup> 4p <sup>2</sup>	As $4s^2 4p^3$	Se 4s <sup>2</sup> 4p <sup>4</sup>	Br 4s <sup>2</sup> 4p <sup>5</sup>
	In 5s <sup>2</sup> 5p <sup>1</sup>	Sn 5s <sup>2</sup> 5p <sup>2</sup>	Sb 5s <sup>2</sup> 5p <sup>3</sup>	Te 5s <sup>2</sup> 5p <sup>4</sup>	I 5s <sup>2</sup> 5p <sup>5</sup>
Char	TI 6s <sup>2</sup> 6p <sup>1</sup>	Pb 6s <sup>2</sup> 6p <sup>2</sup>	Bi 6s <sup>2</sup> 6p <sup>3</sup>	Po 6s <sup>2</sup> 6p <sup>4</sup>	At 6s <sup>2</sup> 6p <sup>5</sup>

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## A Text Book Of

## **INORGANIC CHEMISTRY**

DSC-3A: Paper I

**FOR** 

**B.Sc. Part-I: Semester-I** 

As Per New Revised NEP - 2020 Pattern Syllabus of Shivaji University, Kolhapur w.e.f. June 2023

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## **Preface**

This book is basically intended for B.Sc.-I students for Semester-I Course of Shivaji University, Kolhapur. This book is written according to new NEP-2020 Pattern syllabus being implemented from June 2023 prescribed by Shivaji University, Kolhapur.

It is our great pleasure to present this book to the students and respected teachers in proper time. The subject matter is presented in simple and lucid language. This book covers all the chapters mentioned in the syllabus. The material is presented in a comprehensive way and the sequence of articles in each chapter helps the students to understand the subject with ease.

Different diagrams and illustrative description is provided to enhance the learning and understanding of the matter by students as well as to enable the teachers to explain the difficult concepts properly. The solved numerical examples, long answer type questions, short answer type questions including multiple type questions are given at the end of each chapter.

We are thankful to Nirali Prakashan, Pune for making us a part of their team of Authors. We thank **Mr. Dineshbhai Furia** and **Mr. Jignesh Furia** for publishing this book.

Last but not the least we are very much indebted to Mr. Virdhaval Shinde (Marketing Executive, Kolhapur District) and Mr. Ashok Nanavare (Marketing Executive, Sangli District) for their nice co-operation. We are very much thankful to Mr. Kiran Kamble (Proof Reading), Mrs. Anjali Muley (Graphic Design) and Mr. Akbar Shaikh for a neat and error free D.T.P. of this book.

We hole that this book will be found useful by students and teachers. We will appreciate any suggestions for the improvement of the book.

**Authors** 

# **Syllabus**

## Unit I: Atomic Structure & Periodicity of Elements (08 Hrs.)

- 1.1 Bohr's Theory of Hydrogen Atom and its Limitations
- 1.2 Wave Particle Duality
- 1.3 Heisenberg's Uncertainty Principle
- 1.4 Quantum Numbers and their Significance
- 1.5 Shapes of s, p and d Atomic Orbitals
- 1.6 Electrons Filling Rules in Various Orbitals: (a) Aufbau's Principle,(b) Hund's Rule of Maximum Multiplicity, (c) Pauli's Exclusion Principle
- 1.7 Electronic Configuration of Elements. Stability of Empty, Half-filled and Completely Filled Orbitals
- 1.8 Periodicity of Elements: General discussion of the following Properties of the Elements with reference to s-block elements:
  - (a) Electronic Configuration, (b) Atomic Radii, (c) Ionic Radii,
  - (d) Ionization Energy, (e) Electron Affinity, (f) Electronegativity,
  - (g) Metallic Characters, (h) Reactivity, (i) Oxidation State, (j) Melting and Boiling Points, (k) Chemical Properties.

## Unit II: Chemical Bonding and Molecular Structures:

## Ionic Bonding

(06 Hrs.)

- 2.1 Types of Chemical Bonds: (a) Ionic Bond, (b) Covalent Bond, (c) Co-ordinate bond, (d) Metallic Bond, (e) Hydrogen Bond, (f) Van der Waal's Force
- 2.2 Definition and Formation of Ionic Bond. General Characteristics of Ionic Bonding
- 2.3 Energetics in Ionic Bond Formation
- 2.4 Born-Haber Cycle for NaCl and its Applications
- 2.5 Fajan's Rule, Applications of Fajan's Rule for
  - (i) Polarising Power and Polarisability
  - (ii) Ionic Character in Covalent Compounds
  - (iii) Bond Moment, Dipole Moment and Percentage Ionic Character

# Unit III : Chemical Bonding and Molecular Structures : Valence Bond Theory (VBT) (05 Hrs.)

- 3.1 VSEPR Theory
- 3.1 Concept of Hybridisation, Different Types of Hybridisation and Geometry of following molecules:
  - (i) Linear Geometry BeCl<sub>2</sub> (sp Hybridisation)
  - (ii) Planar Trigonal Geometry BF<sub>3</sub> (sp<sup>2</sup> Hybridisation)
  - (iii) Tetrahedral Geometry SiCl<sub>4</sub> (sp<sup>3</sup> Hybridisation)
  - (iv) Trigonal Bipyramidal Geometry PCI<sub>5</sub> (sp<sup>3</sup>d Hybridisation)
  - (v) Octahedral Geometry SF<sub>6</sub> (sp<sup>3</sup>d<sup>2</sup> Hybridisation)
  - (vi) Pentagonal Bipyramidal Geometry IF<sub>7</sub> (sp<sup>3</sup>d<sup>3</sup> Hybridisation)

## **Unit IV: Acids and Bases**

(04 Hrs.)

- 4.1 Theories of Acids and Bases Arrhenius concept, Bronsted Lowry concept, Lewis concept, Lux-Flood concept. (Definition and Examples only).
- 4.2 Hard and Soft Acids and Bases (HSAB concept)
  - 4.2.1 Classification of Acids and Bases as Hard Soft and Borderline
  - 4.2.2 Pearson's HSAB concept
  - 4.2.3 Acid Base Strength and Hardness-Softness
  - 4.2.4 Applications and Limitations of HSAB concept

## Unit V: p-Block Elements (Group 13, 14, 15)

(07 Hrs.)

- 5.1 Position of Elements in Periodic Table
- 5.2 Characteristics of Group 13, 14 and 15 Elements with Special Reference to Electronic Configuration and Periodic Properties.
- 5.3 Compounds of Group 13, 14 and 15 Elements.
  - 5.3.1 Boron Diborane (only Structure).
  - 5.3.2 Allotropes of Carbon and Phosphorus.
  - 5.3.3 Oxyacids of Nitrogen (HNO<sub>2</sub>, HNO<sub>3</sub>).

# **Contents**

1.	Atomic Structure and	
	Periodicity of Elements	1.1 - 1.34
		,
2.	Ionic Bonding	2.1 - 2.22
3.	Valence Rend Theory (VDT)	
٥.	Valence Bond Theory (VBT)	3.1 - 3.28
4.	Acids and Bases	4.1 - 4.32
		7102
		,
5.	p-Block Elements (Group 13, 14, 15)	5.1 - 5.32
	Model Question Papers	P.1 - P.4

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